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| **Science 8**  **Bohr Models** | Name:  Date:  Block: |

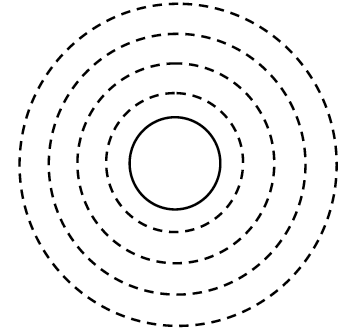
Bohr models show all electrons in each electron shell.

Electrons are distributed in the following way:

* The first electron shell holds a maximum of \_\_\_\_\_\_ electrons.
* Each of the next shells holds a maximum of \_\_\_\_\_\_ electrons.
* Shells cannot be created until the lower shell is \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ filled.

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| In the Bohr model diagram, make sure to:   1. Write the number of \_\_\_\_\_\_\_\_\_ and \_\_\_\_\_\_\_\_\_ in the center of the diagram. 2. Place electrons one at a time before they are paired up. |

Example: Sodium



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| 1. Carbon | 2. Hydrogen | 3. Oxygen |
|  |  |  |
| 4. Chlorine | 5. Potassium | 6. Neon |
|  |  |  |
| 7. Aluminum | 8. Magnesium | 9. Helium |
|  |  |  |
| 10. Boron | 11. Silicon | 12. Beryllium |
|  |  |  |
| 13. Calcium | 14. Phosphorus | 15. Argon |
|  |  |  |
| 16. Sulfur | 17. Nitrogen | 18. Fluorine |
|  |  |  |