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| **Percent Composition** | Name: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_Date: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ |

Find the percent composition by mass the **bold atom** in the following compounds:

1. **Carbon** dioxide - CO2 MM = 44.01g/mol C = 27.29%
2. **K2**CO3 - MM = 138.21 g/mol k = 56.58%
3. Ammonium phosphate – **(Hydrogen) -** (NH4)3PO4 – MM = 149.12 g/mol H =8.13 % (3 SIGS!)
4. C8**H18** - MM = 114.26 g/mol H = 15.91%
5. **C4**H10O - MM = 74.14 g/mol C = 64.80%
6. Mg(N**O3**)2 - MM = 148.33 g/mol O = 64.72%
7. Ca**Cl2 -** MM = 110.98 g/mol Cl = 63.89%
8. Potassium permanganate – **(Oxygen) -** KMnO4 MM = 158.04 g/mol O = 40.50%
9. Ba**F2** - MM = 175.33 g/mol F 21.67%
10. **Calcium** oxalate - CaC2O4 MM 128.10 g/mol Ca = 31.28%

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