**Solutions Chem:**

**Dilutions: Quiz 4b**

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Answer the following:

1. Write dissociation equations to represent the equilibrium present for a saturated solution of each ionic compound.
	1. CaCl2 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
	2. (NH4)3PO4 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
2. If 3.78 L of 0.960 M sodium fluoride solution is added to 6.36 L of 0.550 M calcium nitrate solution, what is the resulting concentration of [Ca+2] and [F-]?

Answers:

1. Write dissociation equations to represent the equilibrium present for a saturated solution of each ionic compound.
	1. CaCl2 CaCl2 (s) 🡪 Ca+2 (aq) + 2Cl- (aq)



* 1. (NH4)3PO4 (NH4)3PO4 (s)🡪 3NH4+ (aq) + PO4-3 (aq)
1. If 3.78 L of 0.960 M sodium fluoride solution is added to 6.36 L of 0.550 M calcium nitrate solution, what is the resulting concentration of [Ca+2] and [F-]?



