**Solutions Chem:**

**Dilutions: Quiz 4a**

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Answer the following:

1. Write dissociation equations to represent the equilibrium present for a saturated solution of each ionic compound.
	1. NaCl \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
	2. Al2O3 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
2. What is the concentration of each ion in the solution formed when 94.5 g of nickel (III) sulphate is dissolved into 850.0 mL of water?

Answers:

1. Write dissociation equations to represent the equilibrium present for a saturated solution of each ionic compound.

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* 1. NaCl NaCl(s) 🡪 Na+(aq) + Cl-(aq)

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* 1. Al2O3 Al2O3(s) 🡪 2Al+3(aq) + 3O-2(aq)
1. What is the concentration of each ion in the solution formed when 94.5 g of nickel (III) sulphate is dissolved into 850.0 mL of water?

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