**Solutions Chem:**

**Polarity and Intermolecular Forces: Quiz 1c**

/5

Answer the following:

1. For each of the following pairs of molecules, determine which is most polar. Explain your reasoning for each:
	1. Carbon Disulfide OR Sulfure Difluoride
	2. Chlorine OR Phosphorus Trichloride
	3. Silicon Tetrabromide OR HCN
	4. Nitrogen Trichloride OR Oxygen Dichloride
	5. Explain your answer for question d.

Answers:

1. For each of the following pairs of molecules, determine which is most polar. Explain your reasoning for each:



1. Carbon Disulfide OR Sulfure Difluoride

CS2 is non-polar

1. Chlorine OR Phosphorus Trichloride

Cl2 is non-polar

1. Silicon Tetrabromide OR HCN

SiBr4 is non-polar

1. Nitrogen Trichloride OR Oxygen Dichloride
2. Both are polar, but OCl2 is MORE polar due to it’s bent structure ☺