**Atoms: Orbital Diagrams: Quiz 2a**

/5

Please show work where necessary!

1. How many electrons can each of the following orbitals hold?

a. 2s = \_\_\_\_\_\_\_\_

 b. 5f = \_\_\_\_\_\_\_\_

1. What is the ***maximum*** number of electrons in the 3rd principal energy level? \_\_
2. Draw the orbitals below and then fill in the orbital diagram for the element **Fe**:



Answers:

1. How many electrons can each of the following orbitals hold?

a. 2s = **\_\_\_\_2\_\_\_\_\_**

 b. 5f = **\_\_\_\_14\_\_\_\_**

1. What is the ***maximum*** number of electrons in the 3rd principal energy level? **\_18\_**
2. Draw the orbitals below and then fill in the orbital diagram for the element **Fe**:

