**Atoms: Atomic Structure and Isotopic Abundance: Quiz 1c**

Please show work where necessary!

/5

1. Completes the following table

|  |  |  |  |  |
| --- | --- | --- | --- | --- |
| **Nuclear Notation** | **# of Protons** | **# of Neutrons** | **# of Electrons** | **Net Charge** |
|  |  |  |  |  |
|  | 75 | 113 | 75 |  |
|  |  |  |  |  |

1. There are 2 isotopes of gallium that occur naturally; Gallium-69 and Gallium-71. The Gallium-69 atoms have a mass of 68.925581 amu and the Gallium-71 atoms have a mass of 70.924707 amu. **What is the percent natural abundance for each isotope?** *Answer to TWO DECIMAL PLACES.*

Answers:



1. Completes the following table

|  |  |  |  |  |
| --- | --- | --- | --- | --- |
| **Nuclear Notation** | **# of Protons** | **# of Neutrons** | **# of Electrons** | **Net Charge** |
|  | 12 | 13 | 10 | +2 |
|  | 75 | 113 | 75 | 0 |
|  | 52 | 74 | 54 | -2 |

1. See Below.

