**Electronic Configurations:**

To find the electronic structure we build from the ground up….

1. **All electrons must be accounted for (strike a single line through a circle for each electron)**
2. **Electrons are filled from the lowest energy to the highest energy**
3. **Two electron per orbital (s has one orbital so 2 electrons, p has three orbitals so 6 electrons)**
4. **When electrons fill orbitals within a subshell each orbital has to be filled by a single electron before an orbital can receive a second electron**



